

Données : masses molaires atomiques :

$$M(\text{H})=1 \text{ g.mol}^{-1}$$

$$M(\text{C})=12 \text{ g.mol}^{-1}$$

$$M(\text{N})=14 \text{ g.mol}^{-1}$$

$$M(\text{O})=16 \text{ g.mol}^{-1}$$

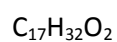
$$M(\text{Na})=23 \text{ g.mol}^{-1}$$

$$M(\text{Cl})=35 \text{ g.mol}^{-1}$$

$$M(\text{P})=31 \text{ g.mol}^{-1}$$

$$M(\text{I})=126 \text{ g.mol}^{-1}$$

Calculer les masses molaires moléculaires des composés suivants :



Corrections

$M(\text{H}_2\text{O}_2) = 2 \times M(\text{H}) + 2 \times M(\text{O}) = 34 \text{ g.mol}^{-1}$	$M(\text{NaCl}) = 58 \text{ g.mol}^{-1}$
$M(\text{CO}_2) = M(\text{C}) + 2 \times M(\text{O}) = 44 \text{ g.mol}^{-1}$	$M(\text{I}_2) = 227 \text{ g.mol}^{-1}$
$M(\text{N}_2) = 2 \times M(\text{N}) = 28 \text{ g.mol}^{-1}$	$M(\text{NH}_3) = 17 \text{ g.mol}^{-1}$
$M(\text{PCl}_5) = M(\text{P}) + 5 \times M(\text{Cl}) = 206 \text{ g.mol}^{-1}$	$M(\text{C}_3\text{H}_9\text{N}) = 59 \text{ g.mol}^{-1}$
$M(\text{NaOH}) = M(\text{Na}) + M(\text{O}) + M(\text{H}) = 40 \text{ g.mol}^{-1}$	$M(\text{C}_{17}\text{H}_{32}\text{O}_2) = 268 \text{ g.mol}^{-1}$
$M(\text{C}_2\text{H}_6\text{O}) = 2 \times M(\text{C}) + 6 \times M(\text{H}) + M(\text{O}) = 46 \text{ g.mol}^{-1}$	$M(\text{NaI}) = 149 \text{ g.mol}^{-1}$